

Ardimerin, a New Dimeric Lactone from the Herb of *Ardisia japonica*

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Abstract – A new ardimerin (**1**) was isolated from the herb of *Ardisia japonica*. Its structure was determined by the elucidation of spectroscopic data and a chemical transformation. Compound (**1**) showed potent radical scavenging effect on DPPH radical (IC₅₀, 0.32 μM).

Keywords – *Ardisia japonica*, ardimerin, DPPH radical scavenging effect

Introduction

Ardisia japonica (Myrsinaceae) is a tropical or subtropical plant, which is widely spread in the seaside area of South Korea, Japan, Taiwan, and China. The fruit and stem of this species have been used as traditional Korean and Chinese medicines to stop cough and uterine bleeding (Perry *et al.*, 1980). Many constituents from *Ardisia* spp. have been reported to have a variety of biological activities such as anti-aging (Takei, 1999), anti-inflammatory (Takei, 1999), skin whitening effect (Kim *et al.*, 2000; Takei, 1999), allelochemical activity (Neal *et al.*, 1998), anti-HIV activity (Piacente *et al.*, 1996), plant growth inhibitory (Lee *et al.*, 1997), 5-lipoxygenase inhibitory (Fukuyama *et al.*, 1993), anti-tubercular (Huang *et al.*, 1981 and 1980), and chitin biodegradative effect (Hirano *et al.*, 1996). In recent years, the role of free radical and reactive oxygen species in human diseases has become more apparent. Thus, the studies on antioxidant activity of compounds derived from natural resources have attracted considerable research interest (Gordon, 1996). In the course of our continuing search for anti-aging agents from medicinal plants, we found that the methanol extract of *Ardisia* spp. exhibited potent antioxidant activity assessed by the scavenging effect of DPPH radical *in vitro*. Bioassay-guided fractionation afforded a unique ardimerin (**1**) with radical scavenging effect on DPPH radical. Herein, we describe the isolation, structural elucidation and radical scavenging activity of **1**.

Materials and Methods

General-UV spectrum was recorded with a HP8453 UV/VIS spectrophotometer. IR spectrum was performed on a Perkin-Elmer model 1750 FT-IR spectrophotometer. FABMS spectra were measured on a JEOL JMX-SX 102 mass spectrometer. High resolution mass measurement was done with a JEOL AX-505H mass spectrometer using m-NBA as a matrix. ESIMS spectrum was also obtained by Mariner (ABI) mass spectrometer. ¹H and ¹³C NMR spectra were recorded in CD₃OD or CDCl₃ at 25°C on a Bruker ARX-400 NMR spectrometer. Chemical shifts (δ) are given relative to TMS, using the solvent peaks [CD₃OD (δ_H 3.30, δ_C 49.0) and CDCl₃ (δ_H 7.26, δ_C 77.1)] as the internal standards. Analytical TLC and preparative TLC were performed on precoated Kieselgel 60 F₂₅₄ plates (0.25 mm, Merck) and Kieselgel 60 F₂₅₄ plates (0.5 mm, Merck), respectively. The silica gel used for column chromatography was Kieselgel 60 (70-230 mesh, Merck).

Plant materials – The whole herb of *Ardisia japonica* was collected in Sunheul, Jeju, Korea. The sample was identified by Professor M. H. Kim of Dep. of Biology, Cheju National University. A voucher specimen (SH203) was deposited at the Laboratory of Natural Products Chemistry, Cheju National University.

DPPH radical scavenging assay – Radical scavenging effects on DPPH (α,α'-diphenylpicrylhydrazyl) radical were determined by the method previously described (Ryu *et al.*, 2001).

Extraction and isolation – The dried whole herb of

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Ardisia japonica (1.0 kg) was extracted with MeOH at room temperature for a day. The MeOH extract (28 g) having DPPH radical scavenging effect (IC_{50} , 26 $\mu\text{g}/\text{mL}$) was partitioned between 60% MeOH and CH_2Cl_2 . The former layer was further partitioned between *n*-BuOH (4.8 g, 17.1%) and H_2O (18 g, 64.3%). The bioactive *n*-BuOH fraction was subjected to an ODS flash chromatography with aqueous MeOH (0, 20, 40, 60, 100%). By the estimation of the DPPH radical scavenging activity of each fraction, the most active 40% MeOH fraction (180 mg) was further fractionated by gel-filtration with MeOH to afford a crude phenolic glycoside fraction (IC_{50} , 1.6 $\mu\text{g}/\text{mL}$). This fraction was finally purified by reversed-phase HPLC with 54% MeOH to yield ardimerin (1, 24 mg, 0.086%). While the latter layer (CH_2Cl_2) was fractionated on ODS flash and silica column chromatography to afford an active fraction. Preparative TLC of this fraction gave four compounds with moderate radical scavenging activity on DPPH radical. Their structural elucidation is in progress.

Compound 1 – Pale brown solid; $[\alpha]_D^{23} +48.2^\circ$ (*c* 0.62, MeOH); UV λ_{max} (log ϵ) 275 nm (0.22), 216 (1.32) and 197 (3.00); IR (film) ν_{max} cm^{-1} 3350 (br, OH), 1710 (C=O), 1605 (aromatic), 1210 (C-O), 960; FABMS (pos) m/z 657 $[\text{M} + \text{H}]^+$, 679 $[\text{M} + \text{Na}]^+$; HRFABMS (pos) m/z $[\text{M} + \text{H}]^+$ 657.4563 (calcd 657.4558 for $\text{C}_{28}\text{H}_{32}\text{O}_{18}$); ^1H NMR (CD_3OD , 400 MHz) and ^{13}C NMR (CD_3OD , 100 MHz) see Table 1.

Acetylation of 1 – A mixture of ardimerin (1, 9.0 mg), Ac_2O (1.0 mL) and pyridine (1.0 mL) was stirred at room temperature overnight. The reagents were evaporated *in vacuo* and the residue was subjected to SiO_2 column chromatography [$\text{CH}_2\text{Cl}_2/\text{MeOH}$ (95:5)] to yield a decaacetylarдимerin (2, 14 mg).

Compound 2 – colorless solid; ESIMS (pos) m/z 1077 $[\text{M} + \text{H}]^+$, 1095 $[\text{M} + \text{H}_2\text{O} + \text{H}]^+$; ^1H NMR (CDCl_3 , 400 MHz) δ 7.72 (1H, s), 5.46 (1H, t, $J = 9.3$ Hz), 5.11 (1H, t, $J = 9.3$ Hz), 4.79 (1H, d, $J = 10.4$ Hz), 4.34 (2H, m), 4.15 (1H, br d, $J = 10.4$ Hz), 3.90 (3H, s, OCH_3), 3.80 (1H, br d, $J = 9.5$ Hz), 2.33 (12H, s), 2.12 (12H, s), 2.08 (12H, s), 2.05 (6H, s); ^{13}C NMR (CDCl_3 , 100 MHz) δ 170.7, 170.3, 169.9, 168.5, 167.9, 161.8, 150.3, 144.5, 141.6, 129.8, 124.3, 118.9, 73.2, 72.4, 68.8, 62.1, 61.8, 21.0, 20.8.

Results and Discussion

Compound 1 was obtained as a pale brown solid. The molecular formula of 1 was established as $\text{C}_{28}\text{H}_{32}\text{O}_{18}$ on the basis of HRFABMS and ^{13}C NMR data. The IR spectrum of 1 revealed absorptions due to a hydroxyl group (3350 cm^{-1}), a carbonyl (1710), an aromatic C=C band (1605), and a C-O band (1210). The ^1H and ^{13}C NMR spectra of 1 (Table 1) were reminiscent of bergenin (4) combined a gallic acid (3) with a *C*-glucose (Ramaiah *et al.*, 1979; Taneyama *et al.*, 1983). A variety of 2D NMR experiments clarified the entire carbon framework of 1 to be listed in Table 1. In particular, HMBC cross peaks between H6 and C1', and between H-1' and H-2' and C6 clarified a β -D-glucosyl residue [δ 81.9, 79.9, 73.8, 72.3, 70.8, 61.2; $J_{\text{H-1'-H-2'}} = 9.8$ Hz and $J_{\text{C1'-H-1'}} = 145.2$ Hz] (Brakta *et al.*, 1993) was directly linked to *meta* position of phenylcarboxylic acid as a *C*-glycoside linkage. No HMBC correlation between the carboxylic carbon and the H-2' of the glucosyl moiety to lead to δ -lactone as showed in bergenin (4) was observed in the HMBC spectrum of 1. A cross peak between 3-OH and C-4 in the HMBC spectrum measured in $\text{DMSO}-d_6$ suggested that a hydroxyl group should be positioned at

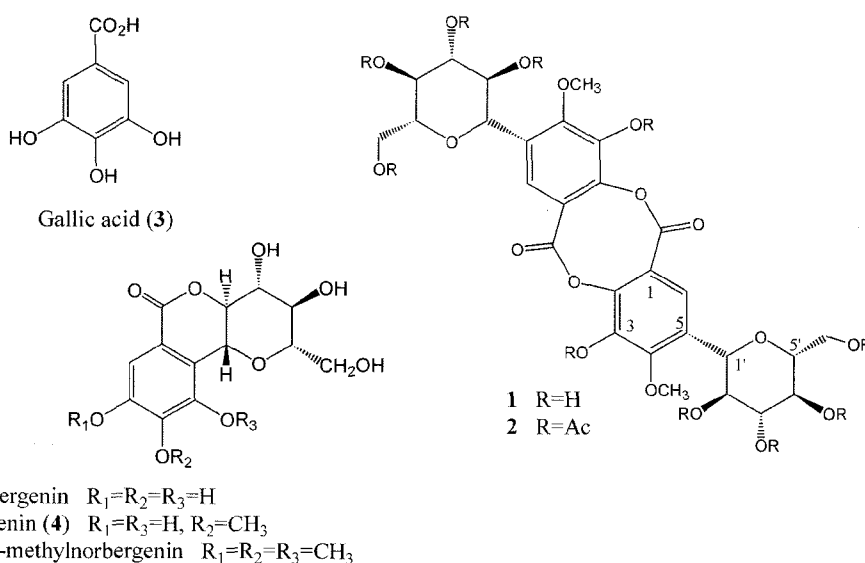


Table 1. ^1H and ^{13}C NMR data of ardimerin (**1**) in CD_3OD at 25°C

Position	δ_{C}	δ_{H} (J Hz)	HMBC
C=O	163.5		
1	118.2		
2	151.1		
3	148.2		
4	140.7		
5	116.0		
6	109.6	6.98 (1H, s)	C=O, C2, C4, C5, C1'
4-OCH ₃	59.9	3.75 (3H, s)	C4
1'	72.3	4.94 (1H, d, 9.8)	C5
2'	79.9	3.98 (1H, t, 9.8)	C5, C3', C1'
3'	73.8	3.65 (1H, m)	C4', C3'
4'	70.8	3.20 (1H, t, 8.9)	C6', C5', C3'
5'	81.9	3.56 (1H, m)	C4', C3'
6'	61.2	3.83 (1H, br d, 11.7)	C5'
		3.45 (1H, dd, 11.7, 2.4)	

C-3. Furthermore, the molecular weight [(M+H)⁺ at *m/z* 1077] of decaacetylarдимerin (**2**) is 748 daltons higher than that of **4**, indicating that **1** is a C₂-symmetric dimer. The carbonyl carbon of one side is connected to 2-OH of the opposite each other, finally forming an eight membered bislactone.

Compound **1** showed potent radical scavenging effect on DPPH radical (IC₅₀, 0.32 μM). Chen's group and Bouchet's group asserted that the radical scavenging activity is dependent on the position of hydroxyl groups (Chen and Ho, 1997; Bouchet *et al.*, 1998). The presence of a second hydroxyl group in the *ortho* or *para* position increases the antioxidant activity due to additional resonance stability and *o*-quinone or *p*-quinone formation. Two hydroxyl groups of **1** orientated to *ortho* each other, while those of bergenin (**4**) to *meta*.

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